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SECTION 3 PROBLEMS Write the answer
on the line to the left. Show all your work
in the space provided. 1. 88% The actual
yield of a reaction is 22 g and the
theoretical yield is 25 g. Calculate the
percentage yield. 2. 6.0 mol of N_2 are
mixed with 12.0 mol of H_2 according to
the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$...

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provided. 1. 88% If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield.

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1. Answers will vary. An example is included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This describes the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

Chapter 9: Standard Review Worksheet

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Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. ____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

CHAPTER 9 REVIEW

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Stoichiometry SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a

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masses in grams of all reactants and products.

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